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Chemistry formulas for iit jee pdf pdf reader full

1. For a general chemical transformation
   $nA + mB \rightarrow pC + qD$
   $\text{Rate} = \frac{-1}{n}\frac{d[A]}{dt} = \frac{-1}{m}\frac{d[B]}{dt} = +\frac{1}{p}\frac{d[C]}{dt} = \frac{1}{q}\frac{d[D]}{dt}$
2. For elementry chemcial reaction
   $mA + nB \longrightarrow$ products
   Theoretical rate $= \frac{dx}{dt} = k[A]^m \times [B]^n$

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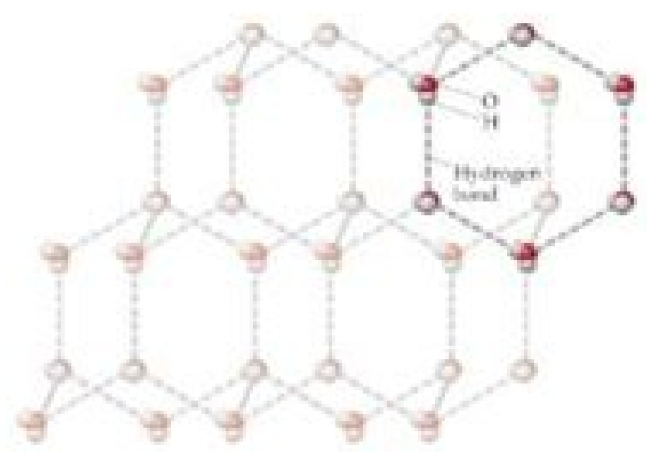
3. For a general reaction: $aA + bB \longrightarrow$ Products
   $\text{Rate} = \frac{dx}{dt} = k[A]^m \times [B]^n$
   Order of reaction w.r.t. A = m overall order of reaction = (m + n)
   Order of reaction w.r.t. B = n
4. Unit of rate constant = $(\text{mole})^{1-n}\ (\text{litre})^{n-1}\ \text{time}^{-1}$
   where, n = order of reaction
5. For a zero order reaction: $A \longrightarrow B$
   $\text{Rate} = \frac{-d[A]}{dt} = k[A]^0 = k$ (constant) $\quad \boxed{K = \frac{[A_0]-[A]}{K}}$
6. For a first order reaction: $A \longrightarrow B$
   $\text{Rate} = \frac{-d[A]}{dt} = k[A]$
   $k = \frac{2.303}{t}\log_{10}\frac{[A]_0}{[A]_t} = \frac{2.303}{t}\log_{10}\left(\frac{a}{a-x}\right)$
7. For a zero order reaction, $t_{1/2} = \frac{[A]_0}{2k}$
   For a first order reaction, $t_{1/2} = \frac{0.693}{k}$
   For an n^{th} order reaction, $t_{1/2} \propto \frac{1}{[A]_0^{n-1}}$ (for $n \geq 2$)

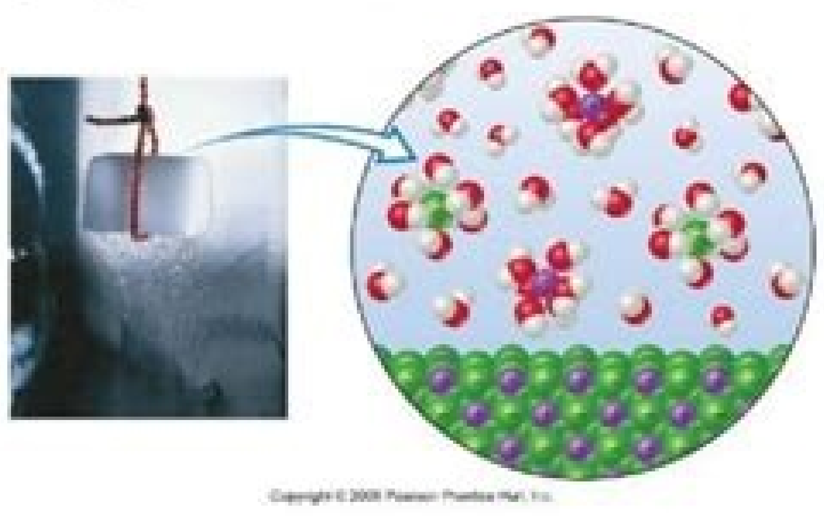
AP Chemistry Chapter 19 Chemical Thermodynamics

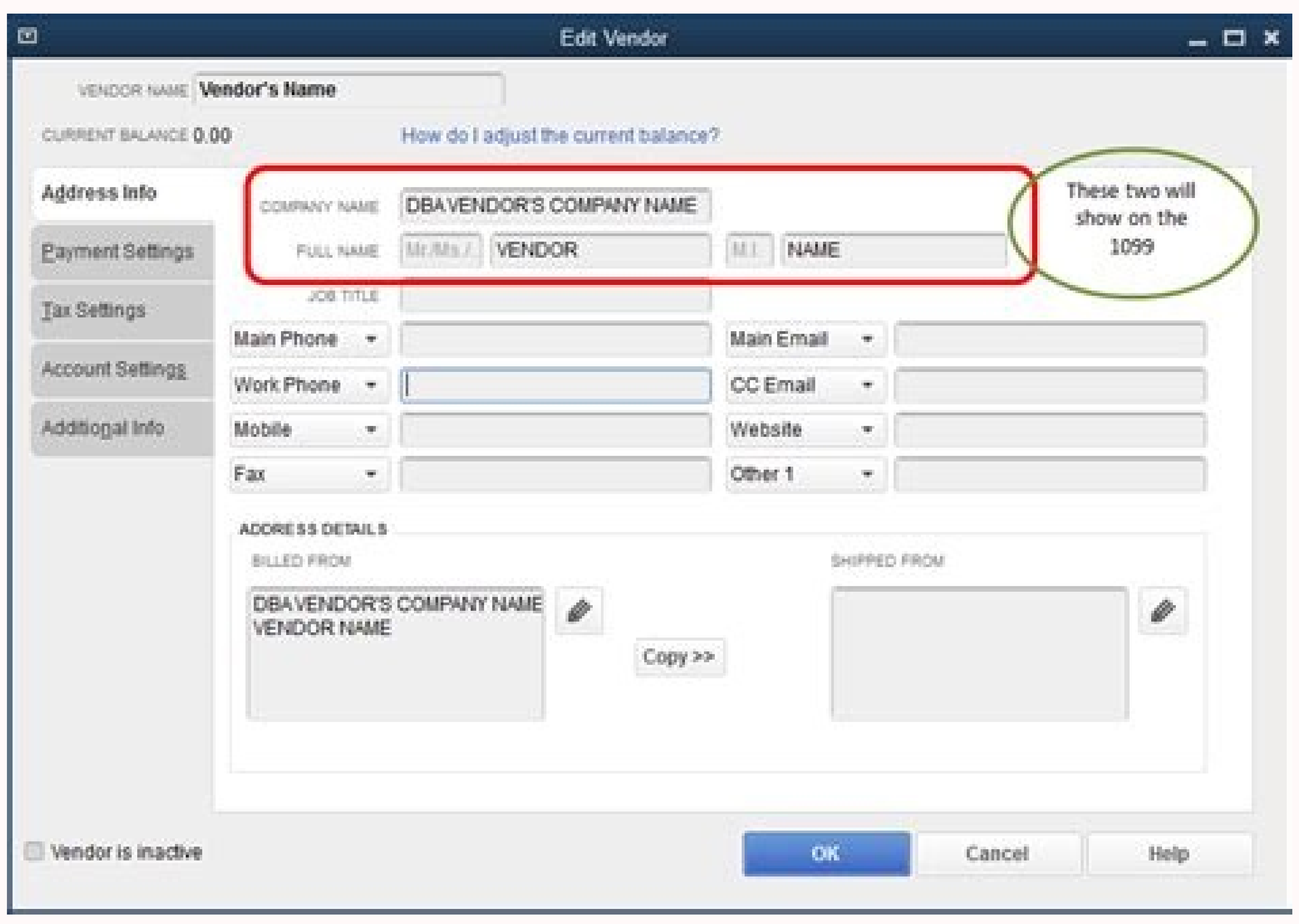
Making Qualitative Predictions About ΔS

- In most cases an increase in the number of microstates (and thus entropy) parallels an increase in:
  - temperature
  - volume
  - number of independently moving particles.
- Consider the melting of ice.
  - In the ice, the molecules are held rigidly in a lattice.
  - The intermolecular attractions in the three dimensional lattice restrict the molecules to vibrational motion only.
  - When it melts, the molecules have more freedom to move (↑ degrees of freedom).
  - The molecules are more randomly distributed.



- Consider a KCl crystal dissolving in water.
  - The solid KCl has ions in a highly ordered arrangement.
  - crystal dissolves → the ions have more freedom → more randomly distributed.
  - However, now the water molecules are more ordered.
  - Some must be used to hydrate the ions.
  - Thus this example involves both ordering and disordering.
  - The disordering usually predominates (for most salts).





(i) O and F are present in the same period of the periodic table. An F atom has one proton and one electron more than O and as an electron is being added to the same shell, the atomic size of F is smaller than that of O. As F contains one proton more than O, its nucleus can attract the incoming electron more strongly in comparison to the nucleus of O atom. Also, F needs only one more electron to attain the stable noble gas configuration. Hence, the electron gain enthalpy of F is more negative than that of O.

(ii) F and Cl belong to the same group of the periodic table. The electron gain enthalpy usually becomes less negative on moving down a group. However, in this case, the value of the electron gain enthalpy of Cl is more negative than that of F. This is because the atomic size of F is smaller than that of Cl. In F, the electron will be added to quantum level $n = 2$, but in Cl, the electron is added to quantum level $n = 3$. Therefore, there are less electron- electron repulsions in Cl and an additional electron can be accommodated easily. Hence, the electron gain enthalpy of Cl is more negative than that of F.

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There is a tough competition in the IIT JEE examination, as there are more than 16.84 lakhs students appearing in this exam and there are around 1.63 lakhs seats including all UG courses like civil, mechanical, electrical, etc. Formula for EMF of cellE0Cell = (standard reduction potential of cathode) - (standard reduction potential of anode) E0 Cell = E0 Cathode - E0 Anode = E0 Right - E0 LeftB. Formula For Elevation of Boiling PointBoiling point elevation = Tb - Tb0Tb = Kb M.m /1000E. Chemistry Formulas and Notes Booklet for Iit Jee p Chemistry Formulas and Notes Booklet for Iit Jee p Amine IIT JEE Organic Chemistry IIT JEE Fluid Mechanics Notes Gyan-Sutra-Chemistry-Formula-Booklet IIT JEE , AIEEE Books for IIT JEE Preparation 0oAULosCnuAC_Math for IIT JEE Choubey.pdf Strategy for IIT JEE Preparation Iit Jee Chemistry Problems With Solutions Bing IIT-JEE 2012 CHEMISTRY practice questions Chapter Wise IIT JEE Notes Mains_Advanced_Phy_Che_Maths Online IIT JEE Study Material for Physics, Chemistry and Maths _ JEE 2016 Toc Solomons Fryhles Organic Chemistry for Iit Jee Bansal Classes Organic Chemistry Study Material for IIT JEE Toc Solomons Fryhles Organic Chemistry for Iit Jee FRICTION NOTES | IIT-JEE | BITSAT | NEET | AIIMS Bansal Classes Chemistry Study Material for IIT JEE It is compulsory to qualify in the IIT JEE examination for getting admission to the undergraduate B. Tech, B.E. or B. Formulas for Chemical EquilibriumA. 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During the last days of preparation, you need to have a quick revision of all important formulas of chemistry. Formulas on ElectrochemistryA. Arch programs at top engineering colleges of India. Formula for Degree of DissociationDegree of dissociation (α) = No. of molecules dissociated/Total no.of molecules taken.Dissociation constant, K(diss):Kdiss = (alpha)2 . Formula For Relative Lowering of Vapour PressureIt is the ratio of lowering of vapour pressure to the vapour pressure of pure solvent.Relative lowering of vapour pressure is given by,= PA0 -p / pa0C. Formula for Reaction Quotient (Q)The reaction quotient has the same algebraic formula as Keq but the current concentrations instead of specifically the equilibrium concentrations are used in calculations.The expression for Q isQ = (C)c (D)d/ (A)a (B)bG. Formula for Concentration of a SolutionConcentration = Quantity of Solute / Volume of Solution B. It is denoted by s and is expressed in mole per litre or gram per litre.Solubility (s) = 1 / Concentration of common ions or number of common ionsD. Formula for Equilibrium Constant in terms of Partial PressureKp = (Pcm1)(Pdm2) /(pan1)(pbn2)Where Kp is equilibrium constant in terms of partial pressure.C. Formula for Equilibrium Constant in terms of Mole FractionKx = XcC XdD / XaA XbBWhere Kx is the equilibrium constant in terms of mole fraction.D. Expression for Relation between Kp and KcKp = Kc (RT)Δnwhere R is gas constant 0.0821 liters.atm/degree/mole and Δn is the total number of molecules of the product – the total number of molecules of the reactants.E. Relation between Equilibrium Constant and Standard Free EnergyΔGo = – RT ln Keqor,ΔGo = – 2.303RT log KeqF. Formula for Raoult's LawThe vapour pressure of a solution is equal to the product of a mole fraction of the solvent and its vapour pressure in the pure state.P / pA = (1-XB)Where xA+xB = 1, therefore xA= 1-xBD. Formula for Ostwald Dilution LawOstwald pointed out that the law of chemical equilibrium is applicable to ionic equilibrium.K = (A)+ (B)- / (AB) = Ca.Ca / C (1-a)Here a denotes alpha sign.B. Formula for Representing Buffer Capacity Buffer Capacity = Number of moles of acid/ Base added to UL of solution/change in pH C. Formula for Gibbs Free Energy ChangeGibbs free energy change in an electrochemical reaction can be expressed as the equivalent of the potential difference.- Del G = nEFD. Law of mass action or chemical equilibriumAt equilibrium Condition,rf = rb, i.e., rate of forward reaction = rate of backward reactionKeq or Kc = (C)x (D)y / (A)a (B)bWhere, Kc = kf / kbWhere the braces and powers are used to indicate "molar concentration".B. Formula for Nernst EquationDependence of electrode potential and EMF on concentration and temperature.For electrode potential:Mn+ + ne- ------- M(s)E = E0 RT In M(s) / nF (Mn+)aqC. Formula for Concentration CellE cell = -(0.0592/2) log (a2/ a1)E0 cell = 04. (Ys)y = Xx.Yy.(S)x+y3. So below are the important formulas of chemistry given topic-wise.Important Formulas of Chemistry are Given below Topic-Wise1. C/ (1- alpha) mol dm-32. Formula for Ionic EquilibriumA. Formula for SolubilitySolubility is the weight of solute in grams present in 100ml of solvent. Every candidate gives their best to get through the IIT JEE examination and you need a unique studying strategy to secure good ranks in the examination. Chemistry is one of the most promising and scoring subjects in the syllabus of IIT JEE. Formula on Solutions and Colligative PropertyA.

The report writing format or steps are as below: 1. Executive summary - highlights of the main report 2. Table of Contents - index page 3. Introduction - origin, essentials of the main subject 4. Body - main report 5. Conclusion - inferences, measures taken, projections 6. Reference - sources of information 7. Appendix Dialogue tags are the phrases like, "he said", "She said'', they attribute the speaker to the dialogue so that the reader always knows who is speaking the dialogue. Dialogue tags are the short lines in a sentence that are used to identify the speaker. The main function of a Dialogue tag in dialogue writing is for identifying who is speaking. Dialogue tags are the phrases like, "he said", "She said'', they attribute the speaker to the dialogue so that the reader always knows who is speaking the dialogue. Dialogue tags are the short lines in a sentence that are used to identify the speaker. The main function of a Dialogue tag in dialogue writing is for identifying who is speaking. Dialogue tags are the phrases like, "he said", "She said'', they attribute the speaker to the dialogue so that the reader always knows who is speaking the dialogue. Dialogue tags are the short lines in a sentence that are used to identify the speaker. The main function of a Dialogue tag in dialogue writing is for identifying who is speaking. The report writing format or steps are as below: 1. Executive summary - highlights of the main report 2. Table of Contents - index page 3. Introduction - origin, essentials of the main subject 4. Body - main report 5. Conclusion - inferences, measures taken, projections 6. Reference - sources of information 7. Appendix The report writing format or steps are as below: 1. Executive summary - highlights of the main report 2. Table of Contents - index page 3. Introduction - origin, essentials of the main subject 4. Body - main report 5. Conclusion - inferences, measures taken, projections 6. Reference - sources of information 7. Appendix

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